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Experiment 19 Freezing Point Depression Answers

calculations

- freezing point depression: $\Delta T_f = K_f \times m_{\text{solute}}$
where $\Delta T_f = T_{\text{final}} - T_{\text{initial}}$
solution 1: $-3.0^\circ\text{C} - 0.5^\circ\text{C} = -3.5^\circ\text{C}$
The freezing point depression of solution 1 is -3.5°C
- $\Delta T_f = T_{\text{final}} - T_{\text{initial}}$
solution 2: $-6.0^\circ\text{C} - 0.5^\circ\text{C} = -6.5^\circ\text{C}$
The freezing point depression of solution 2 is -6.5°C
- Molality of solution 1 (using $K_f = -1.86^\circ\text{C}\cdot\text{kg/mol}$)
 $\Delta T_f = K_f \cdot m_{\text{solute}}$
 $(-3.0^\circ\text{C} - 0.5^\circ\text{C}) = (-1.86^\circ\text{C}\cdot\text{kg/mol})(m)$
 $-3.5^\circ\text{C} = (-1.86^\circ\text{C}\cdot\text{kg/mol})m$
 $-1.86^\circ\text{C}\cdot\text{kg/mol} \cdot m = -3.5^\circ\text{C}$
 $m = 1.88 \frac{\text{mol}}{\text{kg}}$ Molality solution 1: $1.88 \frac{\text{mol}}{\text{kg}}$
- Molality of solution 2 ($K_f = -1.86^\circ\text{C}\cdot\text{kg/mol}$)
 $(-6.0^\circ\text{C} - 0.5^\circ\text{C}) = (-1.86^\circ\text{C}\cdot\text{kg/mol})(m)$
 $-6.5^\circ\text{C} = (-1.86^\circ\text{C}\cdot\text{kg/mol})m$
 $-1.86^\circ\text{C}\cdot\text{kg/mol} \cdot m = -6.5^\circ\text{C}$
 $m = 3.49 \frac{\text{mol}}{\text{kg}}$ Molality solution 2: $3.49 \frac{\text{mol}}{\text{kg}}$
- # of grams of antifreeze per 1000 grams solvent
mass of 250 mL beaker w/ antifreeze - mass of 250 mL beaker = mass of antifreeze
Set proportion as
 $\frac{\text{mass of antifreeze}}{\text{grams solvent}} = \frac{x \text{ mass of antifreeze}}{1000 \text{ grams solvent}}$

A. solution 1
 $(117.70\text{g} - 112.70\text{g}) = \frac{x \text{ mass of antifreeze}}{1000 \text{ grams solvent}}$

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The purpose of this experiment is to use the freezing point depression of the solute ... (19). Notice that the Kf constant only depends on properties of the ... Freezing point depression osmometers are the industry-preferred solution in ... own vapor pressure – a major drawback in the clinical chemistry lab setting.. 12.4.11 Partial answers to the principal questions have now been obtained ... an NaCl solution of the same freezing point depression , using the published ...

In the depression of freezing point experiment , it is found that 1. the vapour pressure of the solution is less than that of pure solvent .. In the depression of freezing point experiment , it is found that the (a) ... (b) have ($4n + 2$) i electrons (c) be planar (d) be cyclic ANSWERS 1 1. A historic point of conflict between Austria Hungary and Russia was over their ... Read and answer the questions for the "19th Century Europe" article: ... (3) Depression of the Freezing Point. (4) Osmotic Pressure ... In an experiment air was drawn successively through a solution of sugar.. Text and lab manual, Lecture final exam. Lecture and ... Experiment 19: Colligative Properties: Freezing Point Depression. 5 ... Never look at the answers first.. NAME, DATE, Chemistry 1^o. Lab 19: Colligative Properties: Freezing-Point Depression and Molar Mass. OBJECTIVES. To become familiar with colligative The melted ice forms a solution with a lower freezing point than that of pure water. Freezing-Point Depression. •The difference in temperature between the determination of molar mass by freezing point depression, determination of a ... chemistry lab experiment 5c answers pdf, advance study assignment experiment. 2238193de0

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